1. A compound contains 26.6% K, 35.4% Cr, and 38.0% O, by weight. Calculate the empirical formula.

2. Two compounds are found to contain Tungsten (W) and Carbon (C). Compound #1 contains 1.84 g W and 0.12 g C. Compound #2 contains 3.69 g W and 0.12 g C. Calculate formulas for each compounds.

3. In a compound containing only carbon and hydrogen, there are 4 g of carbon present for every 1 g of hydrogen. (a) What is the empirical formula for this substance? (b) if the molecular weight of this substance is 30.00, what must the molecular formula be for this substance.

4. N and O atoms form five compounds, each has the definite composition below. Find formulas and name each compound.

Compound	N	O	Formula	Name
1	1.00 g	0.572 g		
2	1.00 g	1.14 g		
3	1.0 g	1.73 g		
4	1.0 g	2.29 g		
5	1.0 g	2.86 g		

5. A compound is found to have a definite composition of C = 27.3% % O = 72.7%. Calculate the formula.

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