

Chapter 8 Problem II

Name _____

Write a balanced chemical equation to represent each of the following chemical reactions:

1. iron + sulfur \rightarrow iron II sulfide

2. zinc + cupric sulfate \rightarrow zinc sulfate + copper

3. silver nitrate + sodium bromide \rightarrow sodium nitrate + silver bromide

4. potassium chlorate \rightarrow potassium chloride + oxygen

5. water \rightarrow hydrogen + oxygen

6. mercury II oxide \rightarrow mercury + oxygen

7. potassium iodide + lead II nitrate \rightarrow lead II iodide + potassium nitrate

8. aluminum + oxygen \rightarrow aluminum oxide

9. magnesium chloride + ammonium nitrate \rightarrow magnesium nitrate + ammonium chloride

10. iron III chloride + ammonium hydroxide \rightarrow iron III hydroxide + ammonium chloride

11. sodium peroxide + water \rightarrow sodium hydroxide + oxygen



12. iron III oxide + carbon \rightarrow iron + carbon monoxide

13. iron + water \rightarrow hydrogen + iron III oxide

14. iron III chloride + potassium hydroxide \rightarrow potassium chloride + iron III hydroxide

15. aluminum + sulfuric acid \rightarrow aluminum sulfate + hydrogen

