Chapter 8 Problem III Name _____

13. Complete and balance these combination reactions.
    a) _____Be + _____O_2 →

    b) _____SO_2 + _____H_2O → _____H_2SO_3

14. Write and balance an equation for the formation of each compound from its elements.
    a) strontium iodide (SrI_2)

    b) magnesium nitride (Mg_3N_2)

15. Complete and balance these decomposition reactions.
    a) _____HI →

    b) _____$Mg(ClO_3)_2$ → _____$MgCl_2$ + _____ _____

16. Write the formula for the binary compound that decomposes to each set of products.
    a) $H_2 + Br_2$

    b) $Na + Cl_2$

17. Complete the equations for these single-replacement reactions that take place in aqueous solution. Balance each equation. If a reaction does not occur write "no reaction."
    a) _____Fe*(s)* + _____$Pb(NO_3)_2$*(aq)* →

    b) _____Cl_2*(g)* + _____NaI*(aq)* →

    c) _____Ca*(s)* + _____H_2O*(l)* →

18. Write the products for thedse double-replacement reactions. Then balance each equations.
    a) _____$NaOH$ + _____$Fe(NO_3)_3$ →

    b) _____$Ba(NO_3)_2$ + _____H_3PO_4 →

19. Write a balanced equation for each reaction.
    a) _____KOH*(aq)* + _____H_3PO_4*(aq)* →

    b) _____H_2SO_4 + _____$Al(OH)_3$ →

20. Write a balanced equation for the complete combustion of each compound.
    a) $HCOOH$

    b) C_7H_{16}

21. Write a balanced equation for the complete combustion of glucose ($C_6H_{12}O_6$).