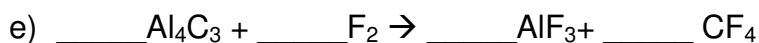
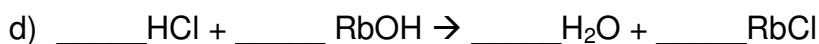
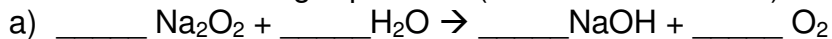
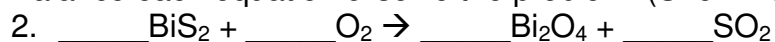


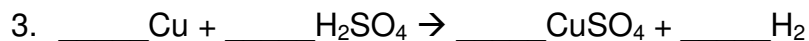
1. Balance the following equations: (Use whole numbers)



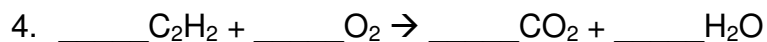
Balance each equation & solve the problem. (Show work)



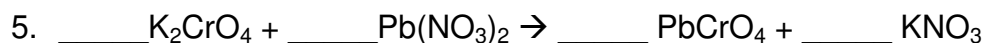
When 1.37 g of Bismuth Sulfide (BiS_2) reacts, what mass of Bi_2O_4 can form? Given excess O_2 .



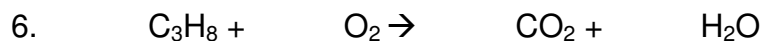
If 0.1280 g of copper dissolve in sulfuric acid, what volume of H_2 gas forms at STP?



To burn 12.00 L of C_2H_2 (acetylene) at STP, _____ L of O_2 gas are needed.



If 3.88 g of K_2CrO_4 are needed to react completely, what mass of $\text{Pb}(\text{NO}_3)_2$ must be used?



To burn 80.0 g of propane (C_3H_8), how many liters of O_2 gas will be needed at STP?