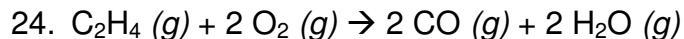
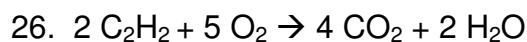
If 2.70 mol C_2H_4 is reacted with 6.30 mol O_2 :

- a) Identify the limiting reagent. _____
b) Calculate the moles of water produced.

If 2.70 mol C_2H_4 is reacted with 6.30 mol O_2 :

- c) Identify the limiting reagent. _____
d) Calculate the moles of water produced.

How many grams of water can be produced by the reaction of 2.40 mol C_2H_2 with 7.4 mol O_2 ?

27. When 84.8 g of iron (III) oxide reacts with an excess of carbon monoxide, 54.3 g of iron is produced. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2 \text{Fe} + 3\text{CO}_2$

What is the percent yield of this reaction.

30. What is the percent yield if 4.65 g of copper is produced when 1.87 g of aluminum reacts with an excess of copper (II) sulfate?



31. What is the difference between an actual yield and a theoretical yield?

Which yield is larger for a given reaction?

How are these values used to determine percent yield?